
Adsorption of Cu^{2+} and Cr^{6+} in Aqueous Solution by a Thermally Modified Biosorbent Based on Cotton Cakes

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Abstract: This work focuses on the retention capacity of Cu^{2+} and Cr^{6+} ions in water by adsorption on a thermally modified biosorbent based on cotton cakes. The solutions treated are synthetic copper nitrate solutions rich in Cu^{2+} ions and potassium dichromate rich in Cr^{6+} ions in distilled water. The adsorption tests were carried out by mixing, with stirring (250 rpm), 0.02 g of the adsorbent for the solution containing the Cu^{2+} ions and 0.04 g for that containing the Cr^{6+} ions with a volume of 50 mL of these solutions. Kinetic studies have shown that the maximum adsorption is obtained after 20 minutes. In addition, the pH of the medium influences the adsorption process and the yield is improved with increasing pH. Amounts of copper and chromium adsorbed increase with increasing initial concentrations of these ions up to 40 mg/L. Similarly, adsorbed amounts increase with increasing mass of the adsorbent. On the other hand, these adsorbed quantities decrease with increasing temperature. The best adopted kinetic models are pseudo-second-order, Intraparticulate and Elovich. The adsorption isotherms conform to the Langmuir and Dubinin-Redushkovic models. The values of thermodynamic parameters such as free enthalpy (ΔH° KJ/K.mol), entropy (ΔS° KJ/K.mol) and isosteric adsorption heat (ΔG° KJ/K. mol) at temperatures of 293; 303; 313 and 323K, all negative, show that we have to do with an exothermic adsorption, physical, ordered and spontaneous. The electrostatic attraction between the Cu^{2+} and Cr^{6+} ions and the adsorbent for the pH values above the pH of the point of zero charges shows that we have to do with a chemical adsorption.

Keywords: Adsorption Isotherms, Cotton Cakes, Copper (II), Chromium (VI), Thermodynamic Parameters

1. Introduction

Water is the central element of all socio-economic processes regardless of the degree of development of society. The increase in agro-industrial activities is putting increasing pressure on the world's freshwater reserves. Indeed, these activities generate a great diversity of chemicals (heavy metals) that flow into the water cycle [1]. Thus, these chemical substances discharged into water are hardly biodegradable and the lack or insufficiency of treatment systems leads to their accumulations in the water cycle [2]. Heavy metals occur naturally in rocks, water, air and soil; they are also produced by human activities, particularly industrial and mining activities [3]. Some metals, such as

zinc, are essential trace elements to humans whose deficiency causes health problems [4]. Others may be toxic, resulting in acute or chronic conditions [3]. This toxicity is highly dependent on the type of element involved in the dose and the duration of the exposure. Metals are released into the atmosphere where they can cover great distances, accumulate on certain sites and contaminate soils, rivers or groundwater [5]. Due to their persistence, heavy metals cannot be reduced or eliminated but are only transformed. This is why their discharge into water and soil is a long-term threat to the environment [6]. This alarming situation should push the actors concerned to find methods of depollution or to propose less polluting industrial processes in order to preserve the water resources and to ensure a continuity of life to all the existing species on earth and also to preserve the fauna and

the flora [7]. It is therefore essential to completely eliminate the heavy metal ions present in the various industrial effluents or reduce their quantity below the allowable thresholds defined by the discharge standards. Several methods have been used, physical and chemical processes have been widely studied to remove heavy metals. Some of these processes are: coagulation [4], flotation [2], chemical precipitation [8], ultrafiltration [9], and electrochemical methods [1]. During the last decades much research has focused on adsorption techniques [10, 4]. This technology has already shown its potential in the treatment inorganic pollutants (heavy metals), toxic organic compounds, natural organic materials and volatile organic compounds. The use of activated carbon has remained for a long time and to this day the most effective adsorbent due to its high adsorption capacity due mainly to its large surface area, but the cost of treatment to obtain it is high, which makes the latter non-competitive from an economic point of view. This has therefore encouraged research by orienting them towards treatment processes using natural materials [11] which are less expensive and widely available such as natural zeolites, volcanic ash, clays and especially the biosorbent such as cotton cakes [12].

2. Material and Methods

2.1. Adsorbent Materials Used

The adsorbent materials used in this work are produced from cotton cakes produced by SODECOTON from Maroua (far north of Cameroon). The choice to use cotton cakes as raw material for the production of the adsorbent is motivated by the valorization of this agricultural waste. It is also obtained at a very favorable cost.

2.2. Adsorbent Production

The manufacture of the adsorbent comprises several steps according to the method of Ibrahim *et al.* (2012) [12]. The raw material once washed, dried and crushed is sieved. The retained particles ($\leq 200 \mu\text{m}$) are impregnated in a phosphoric acid solution with concentrations ranging from 85% (1N) and a mass ratio of 2 g of acid per g of cake. The impregnated particles are oven-dried for 12 hours and are carbonized in a preheated oven at a suitable temperature for a minimum of 2 hours before the start of the experiment to obtain a steady state temperature. The carbonization is carried out at a temperature of 550°C and at a time of 2 hours in order to improve its adsorbent capacity. The adsorbent obtained is cooled in a desiccator. In order to remove any carbonization residues, the adsorbent is washed several times with distilled water to a constant pH. The adsorbent thus washed is dried at 105°C for at least 8 hours in an oven, then cooled in a desiccator and kept airtight in sealed vials until the characterization tests.

2.3. Stock Solution of Cu^{2+} Ions

The stock solution of Cu^{2+} ions with a mass concentration

of 1g/L is obtained by dissolving 0.975 g of copper nitrate ($\text{Cu}(\text{NO}_3)_2$) with a molar mass $M = 241.6 \text{ g/mol}$ at 98% by mass of purity in a liter of 'distilled water.

2.4. Stock Solution of Cr^{6+} Ions

The stock solution of Cr^{6+} ions is prepared using 2.829g of potassium dichromate ($\text{K}_2\text{Cr}_2\text{O}_7$) were dissolved in one liter of distilled water. The standard daughter solutions of chromium in mg/L are obtained by dilution of the stock solution.

2.5. Acid Solution of Diphenylcarbazide

The acid solution of diphenylcarbazide is obtained using the method of Ikhlass *et al.* (2012) [13]. In fact, 0.025g of diphenylcarbazide is dissolved in 30 mL of ethyl alcohol. The resulting solution is transferred to 170 mL of 10% sulfuric acid.

2.6. Characterization of the Adsorbent

The characterization of the adsorbent requires the use of several spectroscopic and volumetric methods. The parameters targeted by this study are: residual moisture, ash content, Iodine index and specific surface.

2.6.1. Residual Moisture

The moisture content or residual moisture is the mass loss after complete desiccation (AFNOR, 1982). It consists of drying a known mass of the sample at 105°C in an isothermal oven to a constant mass (about 24 hours). After cooling in a desiccator, the samples are weighed again. The water content represents the difference in mass before and after the oven.

The water content (TE) as a percentage of wet matter is determined by:

$$TE (\%) = \frac{m_1 - m_2}{m_1 - m_0} \times 100 \quad (1)$$

With m_0 : mass of the empty capsule, m_1 : mass of the capsule containing the wet sample and m_2 : mass of the capsule containing the sample after steaming at 105°C .

2.6.2. Content of Raw Ash

The ash contents are calcination residues at 550°C of the organic materials. They are obtained according to the method of AFNOR (1982). The principle consists in incinerating a test portion after steaming at 105°C , in a muffle furnace at 550°C and weighing the residue obtained. For this, in a muffle furnace, the porcelain capsule containing the sample from the steaming at 105°C was placed at 550°C and was incinerated to white ash free carbon.

The ash content is calculated by the following formula:

$$TC (\%) = \frac{m_3 - m_0}{m_1} \quad (2)$$

With m_0 : mass of the empty capsule (tare), m_3 : mass of the capsule plus ash and m_1 : mass of the sample.

2.6.3. Iodine Index

The iodine number of the adsorbent was determined according to the protocol applied by Ibrahim et al. (2012) [12] or a known mass (g) of the dry adsorbent was mixed with a determined volume of the iodine solution in a vial. After stirring, the mixture was filtered. The filtrate was titrated with sodium thiosulfate in the presence of a few drops of starch.

The calculation of the iodine index, expressed in mg/g, is done using the following equation:

$$\text{Iodine index} = \frac{(V_2 - V_1) \times 0.1 \times MI_2 \times 1.5}{m_A} \quad (3)$$

With V_2 : the volume of thiosulfate titrated to a blank (mL), V_1 : the volume of thiosulphate used to determine the iodine solution after adsorption (mL), MI_2 : the molar mass of iodine (253.81 g/mol) and m_A : the mass of the adsorbent (g).

2.6.4. Specific Surface

The estimation of the specific surface area of an adsorbent is conventionally based on measurements of the adsorption capacity of this adsorbent for a given solute, the molecule of this solute having to have an acceptable surface. It is sufficient for this purpose to determine the value of the adsorption capacity of the monolayer.

For the determination of M_{BM} specific surfaces (m^2/g), we start with the relation:

$$S_{BM} = \frac{Q_{max} \times S \times NA}{M_{BM}} \quad (4)$$

With Q_{max} : ultimate adsorption capacity (mg/g), S_{BM} : the area occupied by a BM molecule (175 Å), NA : Avogadro number and M_{BM} : molar mass of methylene blue (319.853 g/mol).

2.7. pH of the Point of Zero Charges (pH_{PZC})

The pH of the point of zero charges (pH_{PCN}) is defined as the pH of the aqueous solution in which the solid exists under a neutral electric potential. It is determined for all adsorbents using the method of Ibrahim et al. (2015) [12]. Solutions at 0.1 mol/L NaCl and pH between 2 and 10. The pH meter voltCraft was used for pH measurement. However, 0.1g of the adsorbent is brought into contact with 20 mL of each of the solutions contained in capped vials. The suspensions are stirred for 24 hours at room temperature. Each solution is then filtered using a filter paper (whatman) and a new pH measurement is performed. By plotting the curve representing the variation of pH ($pH_f - pH_i$) = f (pH_i). The pH_{PCN} then corresponds to the pH of the solution for which the curve crosses the first bisector.

3. Results and Discussion

3.1. Residual Moisture Values, Raw Ash Content, Iodine Index and Specific Surface Characterizing the Adsorbent

The knowledge of the characteristics of the adsorbent is

very important because it allows to better appreciate its ability to adsorb.

The water content or residual moisture is associated with the amount of water adsorbed by the material. Its value depends on the texture of the material. In this work, a value of 4.248% is obtained (Table 1). This low value of the moisture content of the biosorbent allows us to deduce that the latter has a microporous structure because micropores adsorb very little water and cause low humidity [3].

The raw ash content makes it possible to determine the mineral part. Table 1 shows a percentage ash content of 48.705%. The latter is due to the calcination of cotton cakes which leads to the decrease in percentages of proteins [12].

Table 1 shows the iodine value of 380.71 mg/g. This high value of iodine index makes it possible to conclude that the adsorbent material has on its surface micropores available for the reception of polluting micromolecules [10].

The result of the specific surface area (Table 1) obtained for the adsorbent shows that the latter has significant exchange surfaces reaching 197.77 m^2/g . This high value of the specific surface makes it possible to conclude that this adsorbent complies with the standards of the specific surface area of the adsorbents and this justifies its good adsorbent capacity [14].

Table 1. Characteristics of the adsorbent.

Parameters	Values
Residual moisture	(4.248 ± 0.517) %
Ash content	(48.705 ± 0.191) %
Iodine Index (mg/g)	380.71
S_{BM} (m^2/g)	197.77

3.2. Determination of the pH of the Point of Zero Charges (pH_{PZC})

The value of the pH_{PZC} gives the ideas on the surface charges that the adsorbent materials develop on their surface. Indeed, if the equilibrium pH of the adsorption of the adsorbates on the surface of the materials is greater than the pH_{PZC} , it is deduced that the adsorbent.

Materials develop on their surfaces negative surface charges and on the other hand if the equilibrium pH of adsorption is less than that of the pH of the point of zero charges, the materials rather develop positive surface charges on their surfaces [12]. For the present adsorbent materials, we obtained a value of 2.5 as the pH_{PZC} value (Figure 1).

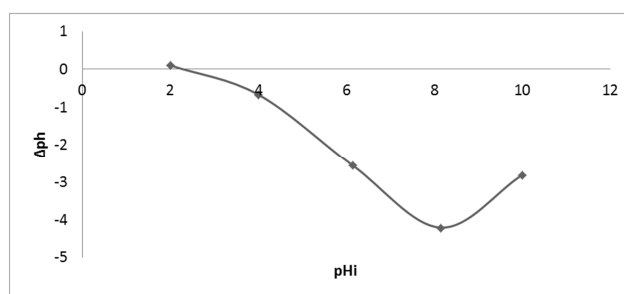


Figure 1. pH of the point of zero charges (pH_{PZC}).

3.3. Kinetics of Adsorption: Contact Time

To measure the stirring kinetics, several identical mixtures at a concentration of 40 mg/L are stirred for different times. We work in a field of agitation duration between 5 to 40 minutes.

The results obtained (Figure 2) show that the retention of the Cu^{2+} and Cr^{6+} ions on the adsorbent is done in two distinct stages: for the first minutes, it is rapid and transient (increasing), then it becomes stable after 20 minutes.

The observed growth is due to the diffusion of the Cu^{2+} and Cr^{6+} ions towards the adsorption sites until reaching an adsorption equilibrium where all the sites become occupied and reveal a similar behavior the Cu^{2+} and Cr^{6+} ions [5].

The observed stabilization is due to the saturation of the adsorption sites [5].

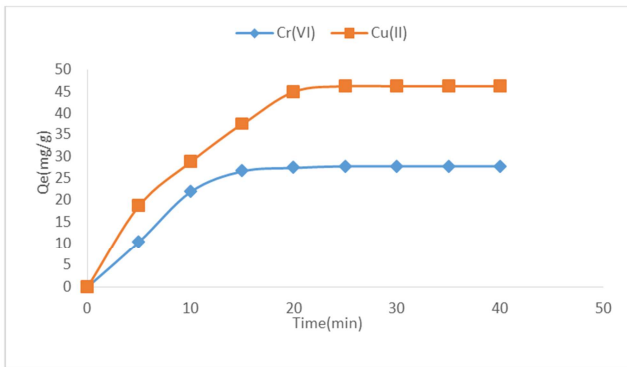


Figure 2. Effect of contact time on adsorption of Cu^{2+} and Cr^{6+} .

3.4. Kinetic Models

The kinetic models studied in this work are those of pseudo first order, pseudo second order, Intraparticle diffusion and Elovich.

3.4.1. Kinetic Model of Pseudo-first Order

The velocity equation of the pseudo-first-order kinetic model is given by the relation:

$$\frac{dq_t}{dt} = k_1(q_e - q_t) \quad (5)$$

Where q_e and q_t are respectively the relative amounts of adsorbed metals at the adsorption equilibrium and at time t , k_1 (min^{-1}) is the pseudo-first order rate constant. The integration of this equation with the initial condition ($t = 0$, $q_e = 0$) leads to the expression:

$$\log(q_e - q_t) = \log q_e - \frac{k_1}{2.303} t \quad (6)$$

Representing the function $\log(q_e - q_t) = f(t)$, we obtain a straight line of slope $-\frac{k_1}{2.303}$ and of ordinate at the origin $\log q_e$. Also, this equation would be suitable for the adsorption of Cu^{2+} and Cr^{6+} ions if the respective correlation coefficient is greater than 0.90 [15]. For this purpose, the values of $\log(q_e - q_t)$ will be calculated and represented as a function of time.

The parameters of the pseudo first-order kinetic model are shown in Table 2.

3.4.2. Kinetic Model of Pseudo-second Order

The pseudo-second-order kinetic model can be represented in the following form:

$$\frac{dq_t}{dt} = k_2(q_e - q_t)^2 \quad (7)$$

Where q_e and q_t are respectively the relative quantities of adsorbed metals at equilibrium and at time t , k_2 is the pseudo-second order rate constant. The integration of this equation with the condition that $t = 0$, $q_0 = 0$ followed by a rearrangement makes it possible to have the following linearized expression:

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{1}{q_e} t \quad (8)$$

By representing the function $t/q_t = f(t)$, the quantity of equilibrium adsorbed metals q_e and the pseudo-second order rate constant k_2 can be determined from the slope and from the ordinate to the origin. The adsorption of Cu^{2+} and Cr^{6+} by the cottonseed cake will follow the pseudo-second order kinetic model if the correlation coefficient of the linear regression of $t/q_t = f(t)$ is greater than 0.90.

Parameters of the pseudo-second-order kinetic model are shown in Table 2.

3.4.3. Kinetic Model of Intraparticle Diffusion

The Intraparticle diffusion rate is determined using the following equation:

$$q_t = k_{int} t^{1/2} + C \quad (9)$$

By representing the function q_t as a function of $t^{1/2}$, we obtain a line of slope k_{int} and an ordinate at the origin C . The equation of the Intraparticle diffusion would be appropriate for the adsorption of the Cu^{2+} and Cr^{6+} ions if the coefficient respective correlation is greater than 0.90.

With k_{int} : constant diffusion rate ($\text{mg/g.h}^{-1/2}$).

The parameters of the kinetic model of Intraparticle scattering are shown in Table 2.

3.4.4. Kinetic Model of Elovich

This model is described by the following differential equation:

$$\frac{dq_t}{dt} = \alpha \exp(\beta q_t) \quad (10)$$

The integration of this equation yields the following linear equation:

$$q_t = \frac{1}{\beta} \ln(\alpha\beta) + \frac{1}{\beta} \ln(t) \quad (11)$$

α ($\text{mg.g}^{-1}.\text{min}^{-1}$) is the initial adsorption rate and β (g.mg^{-1}) is the desorption constant related to the extent of surface coverage and energy of activation for chemisorption.

The kinetic model parameters of Elovich are shown in Table 2.

According to Table 2, all the correlation coefficients R^2 of the pseudo-first-order model are small (less than 0.90) compared to those of the pseudo-second-order models, Intraparticle and Elovich, all of which are very close to the unit for our two metal ions studied.

These results make it possible to conclude that the pseudo-

first-order model is not applicable for the explanation of the elimination of these ions by the adsorbent. However, the values of the R^2 coefficients of the pseudo-second-order model, Intraparticle and Elovich are all greater than 0.90. These results show that only the latter models are applicable to our experimental data.

Table 2. kinetic constants of the four models.

Ions	Pseudo First order			Pseudo 2 nd order			Intraparticle			Elovich		
	q_{max} (mg/g)	K_1	R^2	q_{max} (mg/g)	K_2	R^2	K_{int} (mg·g ⁻¹ ·min ^{-1/2})	C (m ² ·S ⁻¹)	R^2	α	β	R^2
Cu ²⁺	14.381	0.136	0.616	52.910	0.003	0.961	7.863	3.142	0.932	30.719	0.0317	0.972
Cr ⁶⁺	3.431	0.031	0.299	28.818	0.027	0.997	4.525	3.640	0.934	21.322	0.0585	0.902

3.5. Effect of Some Parameters on the Elimination of Cu²⁺ and Cr⁶⁺ Ions

3.5.1. Effect of pH on the Elimination of Cu²⁺ and Cr⁶⁺ Ions

The pH of the aqueous solution is a very important variable that controls the adsorption of an ion at the solid-liquid interface. Its value must be compatible with the ion to be adsorbed. The pH values are measured using a voltcraft-type pH meter equipped with a combined glass-calomel electrode. The pH is set by adding small predetermined amounts of 0.1M sodium hydroxide or 0.1M hydrochloric acid. We will work in a pH range between 2 to 12.

For the pH values of 2 and 4 (Figure 3), the amount of adsorbed Cu²⁺ and Cr⁶⁺ ions are almost constant and varies little with respect to the amount adsorbed for pH 6, 8, 10 and 12 for Cu²⁺ and Cr⁶⁺ ions.

The small variation of the adsorbed quantities of Cu²⁺ and Cr⁶⁺ ions observed is justified by the fact that in strongly acid medium, our adsorbent has developed positive charges on its surface. The surface charges thus developed are therefore not sufficiently suitable for cation adsorption such as Cu²⁺ and Cr⁶⁺ ions. In addition, for pH values in strongly acidic medium there is competition between these ions and H⁺ on the adsorption sites. The presence of the (H⁺) cations is unfavorable for the adsorption of these metal ions because their movements in solution are slowed down by all the hydronium groups which compete with it (electrostatic repulsion of cation and H⁺ ions).

The jump observed for pH values of 6, 8, 10 and 12 for Cu²⁺ ions and pH 8, 10 and 12 for Cr⁶⁺ ions is attributed to an increase in the number of negative surface charges on the surface of the adsorbent following an interaction between Cu²⁺ ions, Cr⁶⁺ and the variable charges that are hydroxyl groups. For these metal ions, the attraction force is increased, therefore the adsorption increases rapidly and this is favorable for the adsorption of all the metal cations because their movement in solution is facilitated by the hydroxyl ions (HO⁻) and all the hydroxyl groups that compete with it [14]. In this case, the interactions of the electrostatic type take on a particular importance (electrostatic attraction of the cations and OH⁻).

At the pH more than 8, the surface is saturated and the metallic ion hydrolyses or form anionic species [8, 13] as shown in figure 3 and table 3.

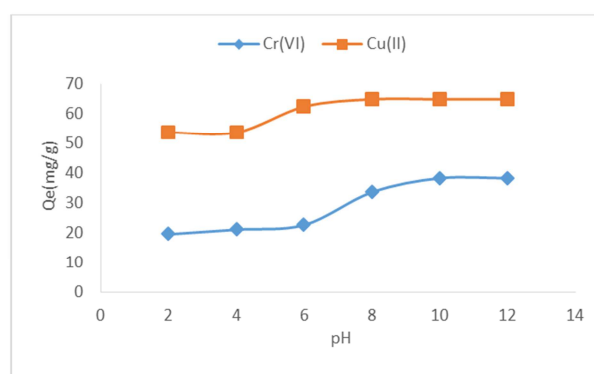


Figure 3. Effect of pH on the removal of Cu²⁺ and Cr⁶⁺ ions.

Table 3. Type of species.

pH	Metallic species	Adsorption
< 4	Cu ²⁺ , Cr ⁶⁺	Weak
[4.5-8]	[CuOH] ⁺ , Cu(OH) ₂ , H ₂ CrO ₄ , [HCrO ₄] ⁻	High
> 8	[Cu(OH) ₂] ²⁻ , [Cu(OH) ₃] ⁻ , CrO ₄ ²⁻	Maximum

3.5.2. Effect of Adsorbent Mass

The effect of the mass of the adsorbent will be studied by contacting different masses ranging from 0.01 to 0.05 g of particle size less than 200 microns. The purpose of this manipulation is to determine the mass of the adsorbent to use for our various manipulations.

Figure 4 shows an increase in the amount adsorbed with the increase in adsorbent mass up to the masses of 0.02 g for Cu²⁺ ions and 0.04 g for Cr⁶⁺ ions.

The increase in the adsorbed quantities of the metal ions with the increase of the mass of the adsorbent is explained by the fact that the more the mass of the adsorbent increases in the medium, the more there is agglomeration of the particles and acceleration diffusion of pollutant molecules into the pores of the adsorbent [16].

This would also be explained by the fact that the smaller the adsorbent mass in the solution (50 mL), the more the medium is not saturated, whereas the diffusion of the metal molecules is not easy [17].

The decrease of the adsorbed quantities observed of the metal ions for the masses of the adsorbent beyond 0.02 g for the Cu²⁺ ions and 0.04 g for the Cr⁶⁺ ions is due to the fact that there are more free ions in solution because the larger the mass is in solution the more the ions are all adsorbed in the solution [16, 17].

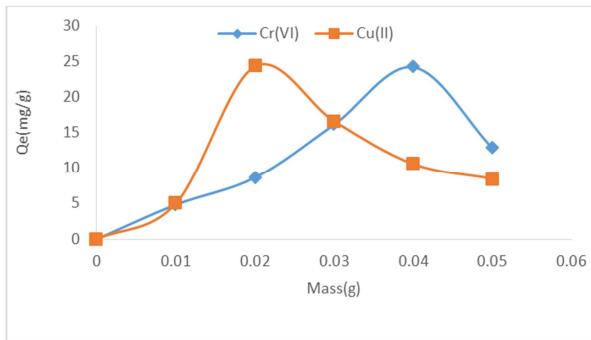


Figure 4. Effect of adsorbent masses on the removal of Cu^{2+} and Cr^{6+} ions in function.

3.5.3. Effect of Initial Concentration of Solutions

To measure the effect of the concentration, several mixtures of different concentrations are stirred at 250 rpm. We will work in a concentration range of 20 to 100 mg/L.

Figure 5 shows an increase in adsorbed quantities of Cu^{2+} and Cr^{6+} ions with increasing concentrations and stabilization at concentrations of 40 mg/L or higher for both ions.

The increase in the amount adsorbed for the first concentrations of the solution observed is justified by the fact that the higher the concentration of the solution, the more molecules of these metal ions are numerous in the solution, the more they are in contact with the solution. Then the adsorbed amount is high [10, 14, 18].

The stabilization of the adsorbed quantities is due to the fact that the adsorption sites or pores are completely saturated, this indicates a gradual saturation of the adsorption sites and the molecules of these ions are free in solution [10, 14, 18].

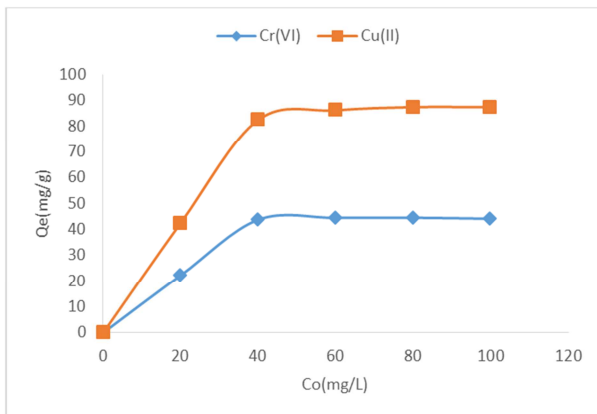


Figure 5. Effect of the initial concentrations of the solutions on the removal of Cu^{2+} and Cr^{6+} ions.

3.5.4. Effect of the Temperature of the Reaction Medium

To measure the effect of temperature, several identical mixtures are stirred at different temperatures. The temperature values are measured in a water bath with a temperature range of 20 to 50°C.

From Figure 6 it is found that the adsorbed amounts of Cu^{2+} and Cr^{6+} ions on the adsorbent decreases as the temperature of the medium increases.

The decrease in the amount adsorbed with the increase in temperature is due to the influence of the pore size. A large fraction of micropores is the same size as the molecule that can enter these pores only under certain temperature conditions [4].

The decrease in the amount adsorbed with the increase in temperature would also be due to a very slow diffusion of the adsorbate and a decrease in active centers available for adsorption. Decreasing the radius of the hydrated ion at high temperature may allow low occupancy of the narrower adsorption sites.

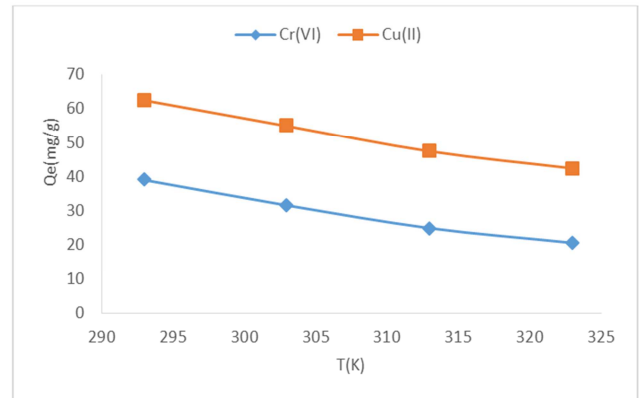


Figure 6. Curve represents the effect of temperature on the removal of Cu^{2+} and Cr^{6+} ions.

3.6. Adsorption Isotherms

The adsorption isotherms of Langmuir, Freundlich, Temkin and Dubinin-Redushkevich are studied in this work [15]. The Langmuir model is widely used to describe adsorption in solution. Freundlich is the second most commonly used model; it applies to many cases, especially in the case of an adsorbent with a heterogeneous adsorption surface. Temkin's assumes that the decrease in heat of adsorption with recovery rates is linear rather than logarithmic.

3.6.1. Insulation Type 1: Langmuir Model

The Langmuir adsorption isotherm is widely used to describe adsorption in solution. The Langmuir isotherm can be represented by the equation:

$$q_e = \frac{ak_L C_e}{1 + k_L C_e} \quad (12)$$

The equation above can be rearranged in the following linear form:

$$\frac{C_e}{q_e} = \frac{1}{ak_L} + \frac{1}{a} C_e \quad (13)$$

With C_e : the residual relative amount of metals at the adsorption equilibrium, q : the relative amount of the adsorbed metal per gram of adsorbent, a : the adsorbed amount at the monolayer and k_L : the adsorption equilibrium constant. The representation of C_e/q_e as a function of C_e makes it possible to determine the coefficients a and k_L for

the adsorption of metals in solution. Thus, adsorption of metals by cotton cake will follow the Langmuir model if the correlation coefficient of the regression of C_e/q_e versus C_e is greater than 0.90.

The parameters of the Langmuir isotherm model are shown in Table 4.

3.6.2. Isotherm Type II: Freundlich Model

The simple and empirical Freundlich model is the second most commonly used model. It is considered that it applies to many cases, in particular in the case of an adsorbent with a heterogeneous adsorption surface:

$$q_e = K_F C_e^{1/n} \quad (14)$$

The linear expression of the Freundlich equation is obtained by taking the logarithm of the equation:

$$\ln q_e = \ln K_F + \frac{1}{n} \ln C_e \quad (15)$$

With K_F and n : the constants respectively reflecting the measurement of the adsorption capacity and the adsorbent-adsorbate affinity. The constant n gives an indication of the intensity of the adsorption. It is generally accepted that low values of n ($0.1 < n < 0.5$) are characteristic of good adsorption, while higher values indicate moderate ($0.5 < n < 1$) or low adsorption. ($n > 1$). The constant " n " is very often replaced by " $1/n$ " or heterogeneity factor.

The parameters of Freundlich's isothermal model are shown in Table 4.

3.6.3. Temkin Isotherm

This model considers a non-uniformity of surface and a preferential occupation of the most adsorbent sites.

The Temkin isothermal model is generally applied in the following form:

$$q_e = \frac{RT}{b} \ln(A C_e) \text{ or } q_e = B \ln A + B \ln C_e \quad (16)$$

With $B = RT/b$, q_e (mg.g^{-1}) and C_e (mg.L^{-1}) are respectively the equilibrium adsorbed amount and the

equilibrium concentration. In addition, T is the temperature (K), R is the gas constant ($8.314 \text{ J.mol}^{-1}.\text{K}^{-1}$). The constant B is related to the heat of adsorption.

The parameters of the Temkin isothermal model are shown in Table 4.

3.6.4. Dubinin-redushkevich Model

Dubinin-Redushkevich's model was inspired by the Dubinin-Polanyi model. The latter is based on the approach of volume filling with the application of Polanyi's potential theory. With this approach, it is considered that the interactions between the adsorbate molecules and the adsorbent surface are determined by a potential field and that the adsorbate volume is only a function of the potential of this field ϵ .

The Dubinin-Redushkevich isotherm is given by the following equation:

$$q_e = q_m \exp(-k\epsilon^2) \quad (17)$$

By linearizing this equation, we obtain the following expression:

$$\ln q_e = \ln q_m - k\epsilon \quad (18)$$

With q_e : quantity of the ion of the adsorbed metal, k : relative constant of the adsorption energy, q_m : theoretical capacity of the micropores, ϵ is the potential of Polanyi ($\epsilon: RT \ln(1+1/C_e)$).

The parameters of the Dubinin-Redushkevich isothermal model are shown in Table 4.

Table 4 presents the values of the R^2 correlation coefficients of the Langmuir, Temkin and Dubinin-Radushkevich isotherms which are all close to unity. Whereas Freundlich's is less than 0.90. In addition, the value of the Freundlich n -value is greater than 1 shows that there is a low affinity between the Cu^{2+} and Cr^{6+} ions and the adsorbent.

These results (Table 4) show that only the Langmuir, Temkin and Dubinin-Radushkevich isothermal models better explain the adsorption phenomenon of Cu^{2+} and Cr^{6+} ions on our adsorbent from cotton cakes.

Table 4. Constants of different adsorption isotherms.

Ions	Langmuir Isotherm			Freundlich Isotherm			Temkin Isotherm			Dubinin-Radushkevich Isotherm		
	a (mg/g)	k_L	R^2	n	k_F	R^2	A (L/g)	B	R^2	K	q_m (mg/g)	R^2
Cu^{2+}	89.285	0.823	0.998	1.17	4.734	0.624	0.171	-107.34	0.996	0.007	7.871	0.994
Cr^{6+}	44.843	1.570	0.999	1.53	4.794	0.546	0.151	-42.059	0.981	0.004	6.26	0.956

4. Thermodynamic Parameters

The adsorbed relative amounts obtained at different temperatures were used to calculate important thermodynamic properties such as Gibbs free standard energy ΔG° (Kj.mol^{-1}), standard enthalpy (ΔH°) (Kj.mol^{-1}) and standard entropy (ΔS°) ($\text{Kj.mol}^{-1}.\text{K}^{-1}$).

The Gibbs free standard energy was evaluated by the following equation:

$$\Delta G^\circ = -RT \ln k_d \quad (19)$$

The equilibrium constant k_d was calculated using the ratio ΔS° :

$$k_d = \frac{q_e}{C_e} \quad (20)$$

With q_e and C_e , the adsorbed amount (mg/g) and the equilibrium concentration of the metals in the solution (mg/L), respectively. The standard enthalpy (ΔH°) and entropy (ΔS°) values were determined from the Van't Hoff equation:

$$\ln k_d = \frac{\Delta S^\circ}{R} = -\frac{\Delta H^\circ}{RT} \quad (21)$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ \quad (22)$$

The representation of $\ln k_d$ as a function of $1/T$ of the Van't Hoff equation is a line of slope $\Delta H^\circ/R$ and intercepts $\Delta S^\circ/R$.

The values of (ΔH°) and (ΔS°) were deduced from the slope and the ordinate at the origin of the equation of the line obtained by the function $\ln k_d = f(1/T)$.

Table 5 shows all the values of the enthalpies of the reaction ΔH° (KJ/K.mol) negative, suggests that the

elimination reaction of Cu^{2+} and Cr^{6+} ions on the adsorbent based on cotton cake is exothermic. The entropies ΔS° (KJ/K.mol) are also all negative, suggests that this chemical reaction leads to the organization of the system (ordered reaction). The free enthalpies ΔG° (KJ/K.mol) of the reaction are also all negative, suggests that this chemical reaction is spontaneous [14, 18].

Table 5. Values of thermodynamic quantities.

T(K)	Cu^{2+}				Cr^{6+}			
	ΔG° (KJ/mol)	ΔH° (KJ/mol)	ΔS° (KJ/mol)	R^2	ΔG° (KJ/mol)	ΔH° (KJ/mol)	ΔS° (KJ/mol)	R^2
293K	-1509				-1587			
303K	-1222				-843			
313K	-922	-9380.68	-26.91	0.995	-255	-18285.8	-57.27	0.980
323K	-716				-99			

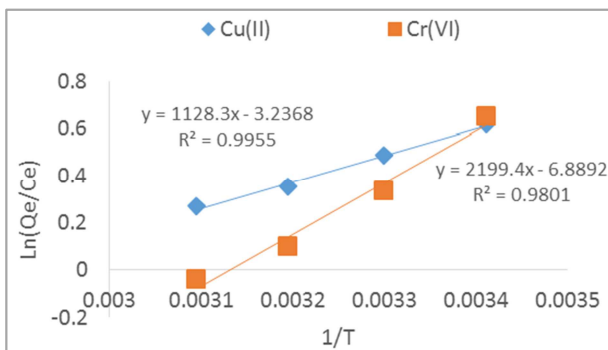


Figure 7. Thermodynamic Parameters.

5. Conclusion

The elimination of Cu^{2+} and Cr^{6+} ions by a thermally modified biosorbent based on cotton cakes is a contribution to water treatment. The adsorption tests were carried out on the synthesized solutions of distilled water by examining the effect of the few reaction parameters such as the contact time, the initial concentration of the solutions, the mass of the adsorbent, the pH of the medium and the temperature of the medium. The results obtained showed that the elimination of these Cu^{2+} and Cr^{6+} ions by the adsorbent produced based on cotton cakes is rapid and the equilibrium is reached at 20 minutes. It is also noted that the adsorbed quantities of these ions increase with the mass of the adsorbent, the initial concentration of the solutions as well as with the pH of the medium. On the other hand, they decrease with increasing temperature. The kinetic models that better express the phenomenon of elimination of Cu^{2+} and Cr^{6+} ions in this work, for which we obtained the correlation coefficients close to unity, are those of pseudo-second order, Intraparticular and Elovich. The adsorption isotherms of Langmuir, Temkin and Dubinin-Redushkevich better express the adsorption mechanism of these two ions on the adsorbent. The low values found for the heat of adsorption (<40 kJ/K.mol) and the electrostatic attraction between Cu^{2+} and Cr^{6+} ions for pH higher than that of the pH of the point of zero charges imply that the adsorption of the ions Cu^{2+} and Cr^{6+} on the adsorbent based on cotton cake is of the physical type (physisorption).

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